

Group Members:

Give the formula for...

- a) **ammonium bromate**
- b) vanadium (V) selenide
- c) lithium nitrite
- d) calcium acetate
- e) **cobalt (III) perbromate**
- f) magnesium sulfide
- g) lithium carbonate
- h) **aluminum hypofluorite**
- i) cobalt(II) sulfide
- j) potassium hydride
- k) cesium peroxide

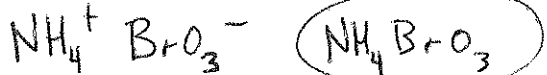
Give the name for...

- a) Ni(OH)_2
- b) NI_3
- c) **HFO_4 (aq)**
- d) $\text{Cr(PO}_4)_2$
- e) N_2O_5
- f) $\text{H}_2\text{S (g)}$
- g) $\text{Fe(NO}_2)_3$
- h) Cl_2O
- i) **H_2SO_4 (aq)**
- j) MnCrO_4
- l) Fe(SCN)_3

Group Members:

Give the formula for...

a) ammonium bromate



b) vanadium (V) selenide



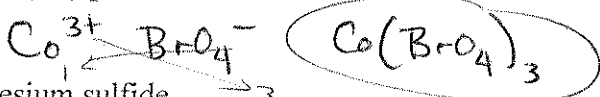
c) lithium nitrite



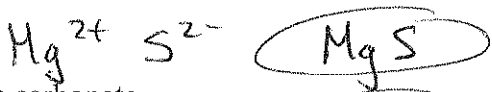
d) calcium acetate



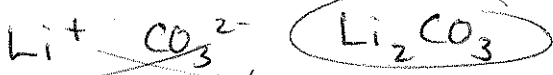
e) cobalt (III) perbromate



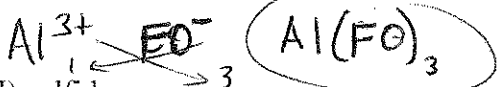
f) magnesium sulfide



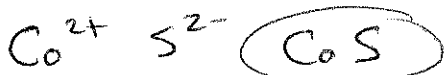
g) lithium carbonate



h) aluminum hypofluorite



i) cobalt(II) sulfide



j) potassium hydride



k) cesium peroxide



Give the name for...



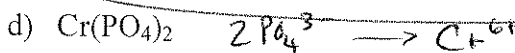
Nickel(II) hydroxide

b) NI_3

nitrogen triiodide

c) $\text{HFO}_4(\text{aq})$

perfluoric acid



chromium(VI) phosphate

e) N_2O_5

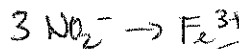
dinitrogen pentoxide

f) $\text{H}_2\text{S}(\text{g})$

hydrogen sulfide as a gas

g) $\text{Fe}(\text{NO}_2)_3$

hydrosulfuric acid
in solution



Iron(III) nitrite

h) Cl_2O

dichlorine monoxide

i) $\text{H}_2\text{SO}_4(\text{aq})$

sulfuric acid

j) MnCrO_4



manganese(I)
chromate

l) $\text{Fe}(\text{SCN})_3$



Iron(III) thiocyanate

Group Members:

Give the formula for...

- a) sodium bromite
- b) vanadium (III) nitrite
- c) lithium phosphide
- d) iron(II) acetate
- e) cobalt (II) hypofluorite
- f) magnesium sulfide
- g) lithium hydrogen sulfite
- h) aluminum perbromate
- i) cobalt(I) sulfide
- j) calcium hydride
- k) magnesium peroxide
- l) titanium(IV) permanganate

Give the name for...

- a) $\text{Ni}(\text{CH}_3\text{COO})_2$
- b) PBr_3
- c) H_3AsO_3 (aq)
- d) $\text{Cr}(\text{PO}_4)_2$
- e) N_2O
- f) H_2Te (g)
- g) $\text{Mn}(\text{CO}_3)_2$
- h) SeF_2
- i) HIO_2 (aq)
- j) BaCr_2O_7
- k) $\text{Fe}(\text{SCN})_3$
- l) AgOH

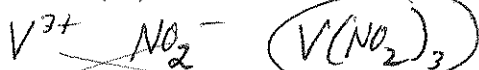
Group Members:

Give the formula for...

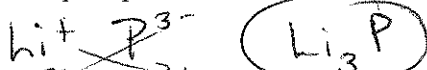
a) sodium bromite



b) vanadium (III) nitrite



c) lithium phosphide



d) iron(II) acetate



e) cobalt (II) hypofluorite



f) magnesium sulfide



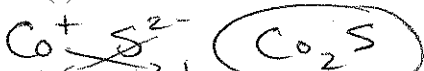
g) lithium hydrogen sulfite



h) aluminum perbromate



i) cobalt(I) sulfide



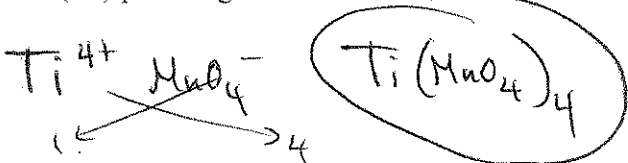
j) calcium hydride



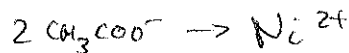
k) magnesium peroxide



l) titanium(IV) permanganate



Give the name for...



a) $\text{Ni}(\text{CH}_3\text{COO})_2$

nickel (II) acetate

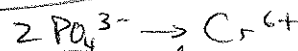
b) PBr_3

phosphorus tribromide

c) H_3AsO_3 (aq)

arsenous acid

d) $\text{Cr}(\text{PO}_4)_2$



chromium (VI) phosphate

e) N_2O

dinitrogen monoxide

f) H_2Te (g)

hydrogen telluride as gas

g) $\text{Mn}(\text{CO}_3)_2$

hydrotelluric acid
in solution



manganese (IV) carbonate

h) SeF_2

selenium difluoride

i) HIO_2 (aq)

iodous acid

j) BaCr_2O_7



barium dichromate

k) $\text{Fe}(\text{SCN})_3$



iron (III)

thiocyanate

l) AgOH

silver hydroxide

NO
roman
numerals
for cation
required

Ag forms only Ag^+
Zn forms only Zn^{2+}
Cd forms only Cd^{2+}